Oxidation numbers (charges) of some common elements.

+1	+2		+3	-3	-2	-1
H^+ Na^+ K^+ Ag^+ Cu^+ Hg^+	Mg^{+2} Sr^{+2} Zn^{+2} Mn^{+2} Cu^{+2} Pb^{+2}	Ca^{+2} Ba^{+2} Fe^{+2} Ni^{+2} Cd^{+2} Hg^{+2}	Al^{+3} Fe^{+3} Cr^{+3} Co^{+3}	N^{-3} (nitride) P^{-3} (phosphide)	O^{-2} (oxide) S^{-2} (sulfide)	F^-(fluoride) Cl^-(chloride) Br^-(bromide) I^-(iodide)

Some important polyatomic ions with their names and oxidation numbers (charges).

+1	-3	-2	-1
NH_4^+ Ammonium H_3O^+ Hydronium	PO_4^{-3} Phosphate PO_3^{-3} Phosphite	SO_4^{-2} Sulfate SO_3^{-2} Sulfite CO_3^{-2} Carbonate CrO_4^{-2} Chromate $Cr_2O_7^{-2}$ Dichromate $C_2O_4^{-2}$ Oxalate MnO_4^{-2} Manganate	OH^- Hydroxide NO_3^- Nitrate NO_2^- Nitrite MnO_4^- Permanganate CN^- Cyanide CH_3COO^- Acetate ClO_4^- Perchlorate ClO_3^- Chlorate ClO_2^- Chlorite ClO^- Hypochlorite

Additionally, in general,

- 1A group elements have always +1 charge, Li^+ ,Na^+, K^+,….
- 2A group elements have always +2 charge, Be^{+2}, Mg^{+2}, Ca^{+2}, Sr^{+2}, Ba^{+2}
- Al^{+3} has always +3 in all compounds
- Hydrogen +1 (with nonmetals) $H^{+1}Cl^{-1}$, -1 (with metals) $K^{+1}H^{-1}$
- Fluorine is the strongest (the most active) nonmetal. So, its charge is always -1
- Oxygen has generally -2 oxidation number . in water H_2O $H^{+1}O^{-2}$ but
  In peroxide it takes -1 oxidation number or O_2^{-2} Na_2O_2 or $Na^{+1}\ O_2^{-2}$
  In OF_2 oxygen takes +2 oxidation number $O^{+2}F^{-1}$
- Nonmetals have generally one negative oxidation number but they have more than one positive oxidation numbers (except F)
  Cl ,Br, I : -1 ,+1, +3 ,+5, +7
  S : -2 ,+4, +6 N : -3 , +1 ,+2, +3, +4, +5
- Some metals have different oxidation numbers
  Cu : +1,+2 Fe : +2,+3 Pb : +2,+4 Mn : +2,+4,+6,+7
- The oxidation number of a monoatomic ion is equal to the charge of this ion.
- The sum of the oxidation numbers for all elements in a compound is zero.